



Synthesis and examination of new PVA-Fe₂O₃-Au hybrid composite

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Abstract

In current work, new hybrid composite have been prepared by mixing gold nanocolloid physically with PVA-Fe₂O₃. Morphological, compositional, structure and optical properties of hybrid composites were studied by transmission electron microscope (TEM), scanning electron microscope (SEM), energy dispersive spectroscopy (EDS), Fourier transform infrared spectroscopy (FTIR), X-ray diffraction (XRD) and UV-visible spectroscopy (UV-VIS). XRD results showed that both Au and Fe₂O₃ have peaks in the structure and that confirmed by EDS and FTIR results. Microscopes results indicated the nanosized nature of prepared materials. UV-vis spectrum show absorption edge at 440 nm which relates to PVA-Fe₂O₃ mixed with Au which means there is a red shift after incorporation of Au in the composites. And the energy gap of composite became equal to 2.48 eV after adding Au solution.

Keywords: Au; Fe₂O₃; nanocomposite; hybrid materials

1. Introduction

Nanoscience and nanotechnology are two captivating sub disciplines within the realm of fundamental science that are currently undergoing significant growth and offer a diverse range of possible applications in the domains of biology, engineering, and medicine. Nanoparticles are widely recognized as substantial synthetic entities within the field of nanotechnology, owing to their enhanced characteristics. The aforementioned characteristics encompass diminutive dimensions, elevated surface area, consistent size dispersion, and a diverse array of potential functions in biological and catalytic processes [1-4].

Polymer nanocomposites, also known as PNC for short, are considered to be among the most promising classes of materials that are now available in the category of advanced materials. Nano reinforced materials display a remarkable combination of qualities, allowing for a wide variety of design possibilities. These composite materials are in high demand in both the academic world and the commercial

world as a result of their tremendous potential and applicability in a wide range of industries [5]. Inorganic semiconductors as filling materials combined with polymers consider hybrid materials give the both good optical and mechanical properties of metal oxides and flexibility of polymers. Many metal oxides that used in fabrication of hybrid materials like TiO₂, CdSe, CdZnS and more [6-9].

In contemporary times, there has been a notable emphasis on the development of magnetic nanoparticles (MNPs), along with a comprehension of their characteristics and the pragmatic implementation of these MNPs in the realm of life science. Magnetic nanoparticles (MNPs) with diverse physicochemical and structural characteristics have been the subject of extensive investigation for a wide range of applications. These applications encompass magnetic drug targeting, hyperthermia, magnetic resonance imaging (MRI), tissue engineering and regeneration, as well as biosensors and bioanalysis [10]. Iron oxide (Fe₂O₃) is a compound that is inexpensive,

favorable to the environment, and resistant to corrosion; as a result, it is a material that is ideal for use in a variety of applications because of its exceptional stability across a wide range of pH and temperature, a substantial absorption of solar light (band gap between 2.1 and 2.2 eV), and magnetic assembling with a high Curie temperature. Fe₂O₃ nanoparticles find application in a wide variety of technological fields, including microwave absorption, catalysis, protection of the environment, gas sensing, magnetic storage, clinical diagnostics and treatment, and many more. Hematite is the form of iron oxide that is the most stable, and it is significant in a wide variety of applications [11-13].

Applications in biomedicine have shown great promise for the use of iron oxide nanoparticles, which have distinctive magnetic characteristics. In more recent times, an expanded research with a variety of iron oxides has been examined as part of a wider study of magnetic particles. Because they readily dissociate under the conditions of an acidic environment, the compounds Fe₂O₃ and Fe₃O₄ are particularly interesting candidates in terms of biocompatibility applications. This is due to the fact that it has been established that these substances are friendly to the human body [14-19].

Owing to their distinctive physicochemical properties, as well as a widespread spectrum of high-activity catalysts, precious metal nanoparticles, such as Au, Ag, and Pd, have gained considerable attention in recent decades. There have been a variety of applications for precious metal nanoparticles so far, including hydrogenation, hydrolysis, fuel cells and electrocatalysis [20]. Mixing Au nanoparticles with Fe₂O₃ in purpose of enhancing the properties both type of nanoparticles intended from many researchers because of the variety of applications especially in medicine and biology as anticancer and antibacterial [21-27].

Herein, the preparation of novel PVA- Fe₂O₃/Au nanocomposites, by synthesizing PVA- Fe₂O₃ nanoparticles through chemical instiu method and mixing with Au nanocolloid by simple solution blending method was attended and the important properties of nanocomposites was tested.

2. EXPERIMENTAL PART

2.1. Preparing of PVA-Fe₂O₃ composite

The preparation and characterization of PVA- Fe₂O₃ nanocomposite had been done according to our published paper [14]. Where the iron chloride (FeCl₃) has been used as iron ions sources, PVA as capping material and sodium hydroxide as an alkaline. 0.1 M FeCl₃ mixed with 1%PVA solution on hot plate after that drops of 1 M NaOH were added slowly to the FeCl₃+PVA solution with continuous stirring at 70 C. brown colloid was presented indicating the formation of PVA-Fe₂O₃ nanocomposite.

2.2. Gold nanoparticles preparation

The synthesis of gold nanoparticles was conducted by a chemical reduction approach, employing gold chloride as the source of gold ions and trisodium citrate as the reducing agent. Where the gold chloride solution was

diluted by dissolving it in non-ionic water to achieve a concentration of 0.005 M. A volume of 5 milliliters of gold solution with a concentration of 0.005 is combined with 35 milliliters of deionized water. The resulting mixture is placed on a magnetic device and heated until it reaches its boiling point. The mixture is then allowed to boil for a duration of 15 to 20 seconds. One milliliter of Sodium citrate, which was made by dissolving 1 gram of TSC powder in 100 milliliters of deionized water, is introduced into the aforementioned mixture. Subsequently, the heat source is deactivated and the mixture is left on the magnetic stirring apparatus. The solution undergoes a transition from coloration to colorlessness, followed by a subsequent change to a dark red hue (specifically, a violet shade that ultimately transforms into a vibrant red tone). This color transformation serves as an indicator for the attainment of gold nanoparticles with dimensions that are comparatively diminutive (less than 20 nm in size).

2.3. PVA-Fe₂O₃/Au nanocomposite preparation

PVA-Fe₂O₃/Au nanocomposite have been prepared by using solution mixing method which it done by mixing 1ml from PVA- Fe₂O₃ with 1 ml of Au NPs together and left on magnetic strerr at room temparature for 5 hours after that we get a homogenous solution.

2.4. Instrumentation

The determination of the crystal structure, surface morphologies, chemical compositions, and optical characteristics of the hybrid nanomaterials was conducted using. The X-ray diffraction (XRD) instrument used in this study was the D2 phaser model manufactured by Bruker, the CuK 1.544 nm target created the X-ray beam, 10 mA at 30 KV. The prepared samples were tested by some of instruments like: TEM Electron Microscopy model a JEOL JEM. 2010 with resolution 0.23 nm and 200kV accelerating voltage(FESEM) model MIRA3 from TE SCAN Company, ALPHA II from Bruker Fourier transformation infrared spectroscopy (FTIR)was used in this investigation, and UV-Visible spectroscopy (UV-Vis) type 1800 from Shemadzu Spectroscopy in the range of wavelength (290-900) nm was employed for the UV-Visible spectroscopy (UV-Vis).

3. Results and discussions

The X-ray diffraction (XRD) procedure has been employed to investigate the crystal structure of the synthesized materials. picture 1 illustrates the diffraction peaks observed in the PVA-Fe₂O₃-Au composite material, while the inset picture displays the patterns specifically observed in the PVA-Fe₂O₃ component. The amorphous nature of the PVA was confirmed by the presence of a broad peak corresponding to the polymer matrix at an angle of $2\theta = 21^\circ$. The diffraction peaks associated with Fe₂O₃ are plainly observable. The confirmation of Fe₂O₃ production was established by the presence of two diffraction peaks at $2\theta = 33.50^\circ$ and $2\theta = 38.00^\circ$. A novel observation was made at a diffraction angle of $2\theta = 27^\circ$, indicating the presence of gold nanoparticles (Au NPs). and this results matching ref.

[28].

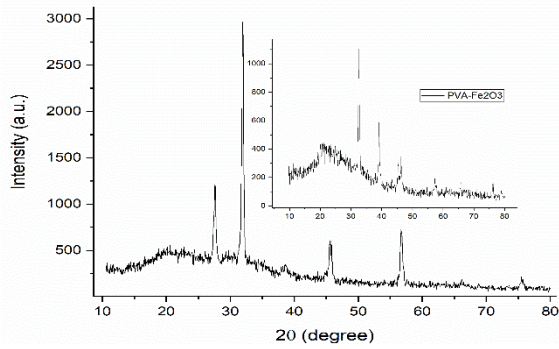


Figure 1. XRD peaks of PVA-Fe₂O₃-Au composite.

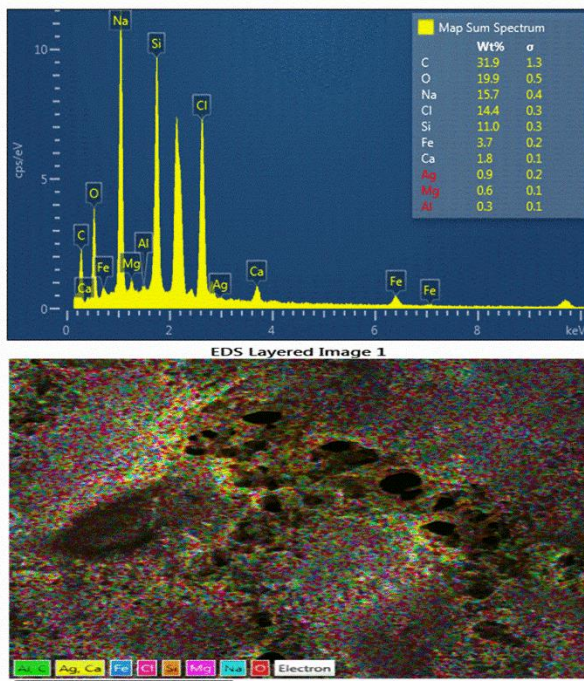


Figure 2. EDS spectrum and mapping of PVA-Fe₂O₃-Au composite.

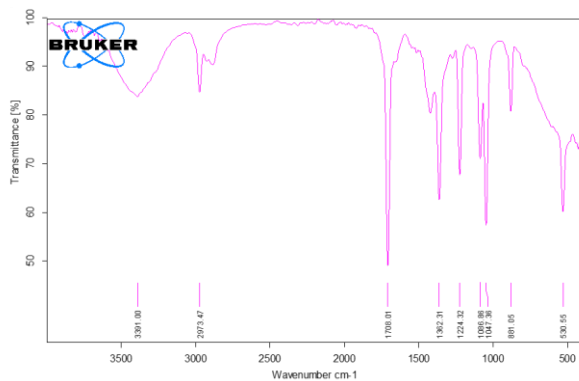


Figure 3. FTIR spectrum of PVA- Fe₂O₃-Au hybrid composite.

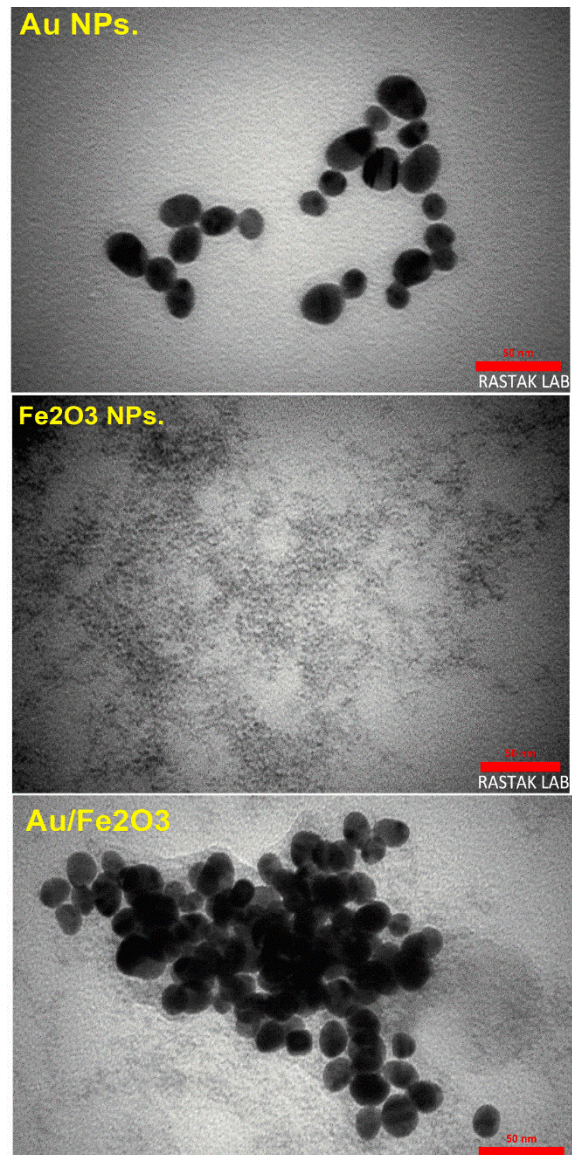


Figure 4. TEM graphs of Au NPs , PVA- Fe₂O₃ and PVA-Fe₂O₃-Au composites .

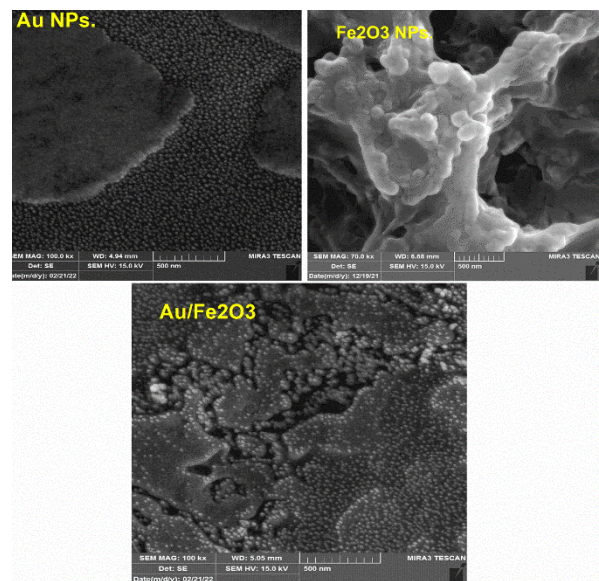


Figure 5. FESEM graphs of Au NPs, PVA- Fe₂O₃ and PVA-Fe₂O₃-Au composites.

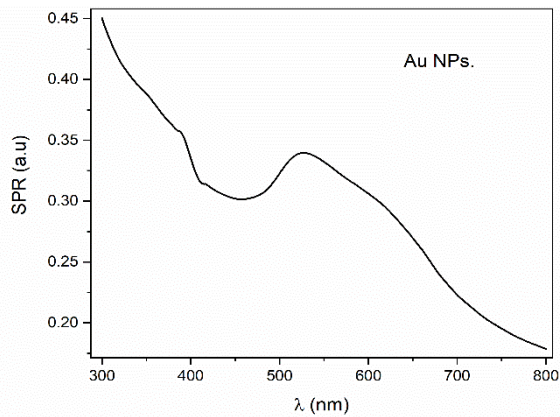


Figure 6. Absorption spectrum and corresponding energy gap of Au NPs.

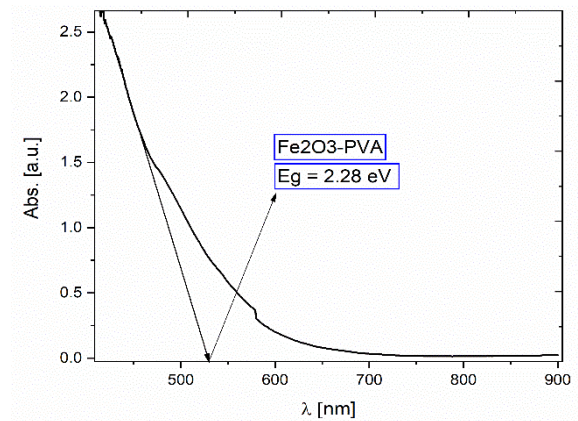


Figure 7. Absorption spectrum and corresponding energy gap of PVA-Fe₂O₃.

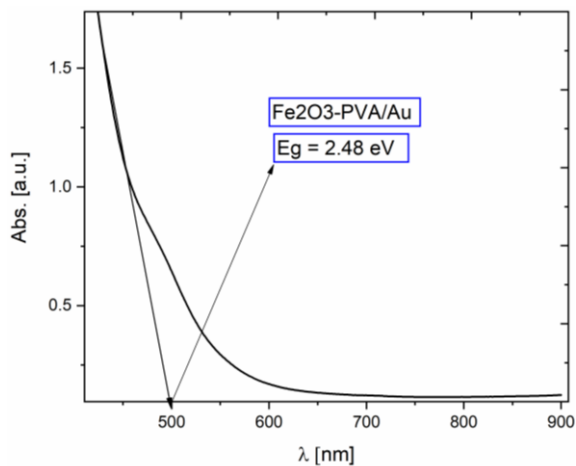


Figure 8. Absorption spectrum and corresponding energy gap of PVA-Fe₂O₃-Au composites.

EDX analysis was performed to ascertain the chemical composition of the prepared samples and identify the components present. Figure 2 presents the EDX spectra and mapping image of the PVA-Fe₂O₃-Au nanocomposite. The graphic illustrates the presence of numerous ingredients within the remedy. and this results matching ref [29].

Figure 3 displays the FTIR spectrum of PVA-Fe₂O₃-Au nanocomposite, where the spectrum has many transmission peaks belonging to PVA, Fe₂O₃, Au and the interaction between all compounds. All the peaks are assigned to the PVA bonds: The O-H stretching vibration of the hydroxyl group, which is sensitive to the hydrogen bond, is responsible for the broad absorption band that can be found at approximately 3391 cm⁻¹ [30,31].

The peak about 532 cm⁻¹ is related to Fe-O stretch of Fe₂O₃ and no peaks belong to Au and that confirm there is no chemical interaction between Fe₂O₃ and Au and the interaction is physical only [32,33].

The morphologies and sizes of the Au nanocolloid were reconnoitered using the TEM which is signified by figure 4a. from the figure, it can be seen that the nanoparticles are dispersed in a good way, as there are no agglomerations, and mono particle can be distinguished from another easily, as well as most of the particles are identical in shape and size. The dominant shape of most

of them is spherical, and their diameters are about 10 nm. The form of the PVA-Fe₂O₃ particles is depicted in figure 4b. This picture illustrates the encapsulation of iron oxide nanoparticles within a polymer matrix, resulting in the formation of irregularly shaped particles. Figure 4c shows the TEM image of mixing both Fe₂O₃ and Au. FESEM examination displays the formation of Au, Fe₂O₃ and their mixing (figure 5). In the case of Au nanoparticles the figure shows formation of spherical particles with nanosized distribution. Through this figure, the polymer envelops the iron oxide nanoparticles, resulting in the formation of irregular particle forms. FESEM measurement exposed that the produced nanocomposites have different sizes and shape and the diameters were in nanoscale. There is a possibility that the smaller nanoparticles aggregated together to form the larger ones and this result agreed with reference [30]. UV-Vis spectroscopy is used to study the absorption spectra and corresponding energy gaps of Au nanoparticles, PVA-Fe₂O₃ and PVA-Fe₂O₃-Au nanocomposites. UV-Vis spectroscopy is used to study the absorption spectra of gold nanoparticles. Figure 6 shows the UV-Vis absorption spectrum of Au NPs. the figure displays the surface plasmon resonance position around 520 nm.

Figure 7 illustrates the absorption spectrum of the PVA-Fe₂O₃ nanocomposite. The absorption edge of the PVA-Fe₂O₃ composite material was observed at approximately 470 nm. The energy gap was determined to be 2.28 electron volts (eV).

Figure 8 represents the UV-Vis spectrum of absorption for PVA-Fe₂O₃-Au nanocomposites. The UV-visible spectrum exhibits an absorption edge at a wavelength of 440 nm, indicating the presence of a nanocomposite structure consisting of PVA-Fe₂O₃ covered with Au. This observation serves as confirmation of the aforementioned materials' composition, which means there is a red shift after incorporation of Au in the composites and this result agreed with results that got from [33] and [34]. the corresponding energy gap was calculated and equal to 2.48 eV.

4. Conclusions

Gold (Au) nanoparticle, PVA-Fe₂O₃ nanoparticle and PVA-Fe₂O₃-Au nanocomposite were synthesized in this work. These prepared composite materials were

characterized by using different analytical methods such as: TEM, FESEM, and UV-Vis spectroscopy. XRD results showed that both Au and Fe₂O₃ have peaks in the structure and that confirmed by EDS and FTIR results. The morphology of the prepared nanocomposite was analyzed by TEM and FESEM, which reveals the creation

of a tiny spherical shape of PVA- Fe₂O₃-Au nanocomposites with a diameter that falls between 40 and 60 nm. there is a red shift absorption edge after incorporation of Au in the composites. And the energy gap of composite became equal to 2.48 eV after adding Au solution.

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